

UIC FILE COPY

1

REPORT DOCUMENTATION PAGE

AD-A211 854

TIC

LECTE

S AUG 3 1989

D

D

4. PERFORMING ORGANIZATION REPORT NUMBER(S) CL-TR-89-0228		1b. RESTRICTIVE MARKINGS	
6a. NAME OF PERFORMING ORGANIZATION Geophysics Laboratory		3. DISTRIBUTION/AVAILABILITY OF REPORT Approved for public release; Distribution unlimited	
6c. ADDRESS (City, State, and ZIP Code) Hanscom AFB Massachusetts 01731-5000		5. MONITORING ORGANIZATION REPORT NUMBER(S)	
8a. NAME OF FUNDING/SPONSORING ORGANIZATION		6b. OFFICE SYMBOL (If applicable) PHK	
8c. ADDRESS (City, State, and ZIP Code)		7a. NAME OF MONITORING ORGANIZATION	
		7b. ADDRESS (City, State, and ZIP Code)	
		9. PROCUREMENT INSTRUMENT IDENTIFICATION NUMBER	
		10. SOURCE OF FUNDING NUMBERS	
		PROGRAM ELEMENT NO. 61102F	PROJECT NO. 2303
		TASK NO. G2	WORK UNIT ACCESSION NO. 01
11. TITLE (Include Security Classification) Spacecraft Interactions as Influenced by Thermochemical Considerations			
12. PERSONAL AUTHOR(S) Edmond Murad			
13a. TYPE OF REPORT Reprint	13b. TIME COVERED FROM _____ TO _____	14. DATE OF REPORT (Year, Month, Day) 1989 August 23	15. PAGE COUNT 6
16. SUPPLEMENTARY NOTATION Reprinted from J. Spacecraft, Vol. 26, No. 3, May-June-1989			
17. COSATI CODES		18. SUBJECT TERMS (Continue on reverse if necessary and identify by block number)	
FIELD	GROUP	SUB-GROUP	Thermochemistry Spacecraft Interactions
19. ABSTRACT (Continue on reverse if necessary and identify by block number) Spacecraft in low Earth orbit undergo harsh environmental interactions that sometimes resemble the phenomenon observed in high-temperature vaporization and oxidation. For example, it has been reported that an osmium surface is removed, and surfaces of aluminum, silver, copper, iridium, niobium, nickel, and tungsten, are oxidized, even though the surface temperature when the measurements are made is -250°K. Similar results are observed for the case of nonmetallic samples, e.g., graphite and synthetic organic materials, which are removed quickly. The purpose of this paper is to examine the data from the space experiments in light of thermochemical data and to consider possible implications or insight that may be gained for planning future experiments.			

89 8 31 010

20. DISTRIBUTION/AVAILABILITY OF ABSTRACT <input type="checkbox"/> UNCLASSIFIED/UNLIMITED <input checked="" type="checkbox"/> SAME AS RPT. <input type="checkbox"/> DTIC USERS		21. ABSTRACT SECURITY CLASSIFICATION UNCLASSIFIED
22a. NAME OF RESPONSIBLE INDIVIDUAL Edmond Murad		22b. TELEPHONE (Include Area Code) (617) 377-3176
22c. OFFICE SYMBOL PHK		

SL-TR-89-0228

Spacecraft Interactions as Influenced by Thermochemical Considerations

Edmond Murad*

U.S. Air Force Geophysics Laboratory, Hanscom Air Force Base, Massachusetts

Spacecraft in low Earth orbit undergo harsh environmental interactions that sometimes resemble the phenomenon observed in high-temperature vaporization and oxidation. For example, it has been reported that an osmium surface is removed, and surfaces of aluminum, silver, copper, iridium, niobium, nickel, and tungsten, are oxidized, even though the surface temperature when the measurements are made is $\sim 250^{\circ}\text{K}$. Similar results are observed for the case of nonmetallic samples, e.g., graphite and synthetic organic materials, which are removed quickly. The purpose of this paper is to examine the data from the space experiments in light of thermochemical data and to consider possible implications or insight that may be gained for planning future experiments.

Nomenclature

ΔF	= free energy of reaction, kcal/mole
$\Delta S(i)$	= entropy of species i , cal deg^{-1} mole^{-1}
$\Delta S_r(i)$	= entropy of reaction for reaction i , cal deg^{-1} mole^{-1}
$\Delta H_f(i)$	= heat or enthalpy of formation of species i , kcal/mole
$\Delta H_v(i)$	= heat or enthalpy of vaporization of species i , kcal/mole
$\Delta H_r(i)$	= heat of reaction for reaction i , kcal/mole
R	= gas constant = 1.987 cal
T	= temperature, $^{\circ}\text{K}$
K_{eq}	= equilibrium constant, which is defined as the product of the partial pressures (or activities of the products divided by the product of the partial pressures or activities of the reactants)
s	= solid phase
c	= condensed phase
g	= gas phase
E_{cm}	= energy of collision in the center of mass frame of reference
E_{Lab}	= energy of collision in the laboratory frame of reference
m_t	= mass of target, amu
m_p	= mass of projectile, amu

Introduction

THE environment of spacecraft in low Earth orbit (200-400 km altitude) is rather harsh because of two important factors: 1) the ambient atmosphere is very corrosive ($\sim 70\%$ is atomic oxygen),¹ and 2) the collisions between the ambient species and spacecraft surfaces are rather violent because of the orbital velocity ($\sim 7.7 \text{ km} \cdot \text{s}^{-1}$). Table 1 summarizes the composition of the atmosphere at altitudes of 250, 300, 350, and 400 km. The orbital velocity of $7.7 \text{ km} \cdot \text{s}^{-1}$ means that any surface directly in the velocity vector will be bombarded by O having an energy of $\sim 4.9 \text{ eV}$ and N₂ having an energy of $\sim 8.6 \text{ eV}$ (these are mean energies; because of the Maxwellian distribution in the velocities, the energies can be somewhat higher). In this paper we will not consider surface catalyzed reactions leading to optical emissions, such as the

Received Feb. 29, 1988; revision received Nov. 29, 1988. This paper is declared a work of the U.S. Government and is not subject to copyright protection in the United States.

*Research Chemist, Manager of Spacecraft Contamination Program. Member AIAA.

Shuttle glow. Such reactions are basically low-temperature phenomenon, and interested readers are referred to a number of reviews in Refs. 2-5.

Experiments

Intentional and unintentional materials experiments on the Space Shuttle have indicated that this interaction between the surfaces and the ambient atmosphere leads to effects that resemble those of high-temperature vaporization and oxidation. For example, Peters et al.⁶ reported the results of a study in which thin films of Ag, C, Os, and Au were exposed to that environment on Shuttle flight STS-4 at 310 km altitude. They observed that Ag gained weight, C and Os disappeared, and Au was unaffected. On another flight, STS-8, which flew at an altitude of 225 km, Peters et al.⁷ reported that Al, Nb, Ni, and W had gained mass, presumably due to the formation of their respective oxides. In addition to the experiments on metals, samples of carbon were eroded quite efficiently⁷; diamond was the least affected of the carbon samples, and graphite was the most affected. The studies on carbon showed the presence of striated valleys in the eroded samples, an observation that is referred to as carpeting. In all of these experiments, the products formed when materials are lost were not identified; on the other hand, mass gain was identified after the flights as being due to the formation of oxide coatings. It should be stressed that the bulk temperature when these experiments were conducted was quite low, between 225 and 300 $^{\circ}\text{K}$, depending on the solar zenith angle.

Experiments on the effect of orbital environment on carbon-containing materials indicate that such materials are even more unstable than the metals discussed previously. The most reactive material is polyethylene and the least reactive material is Teflon.⁸ As in the case of the metals, there is no information as to the nature of the products that are formed in these reactions. Undoubtedly, CO and CO₂ are formed.

The observations may be shown schematically in Fig. 1. Table 2 shows a summary of the observations as reported in a recent review,⁹ as well as a listing of the thermal conductivity of the elements as reported in a review.¹⁰

Discussion

In this analysis, we will follow the tree shown in Fig. 1 and begin by discussing the observations on 1) metals (including Si) and 2) nonmetals (carbonaceous materials).

Before we consider the experiments, however, it will be worthwhile to define the terms and units that will be used in



odes

for

1A-1 20

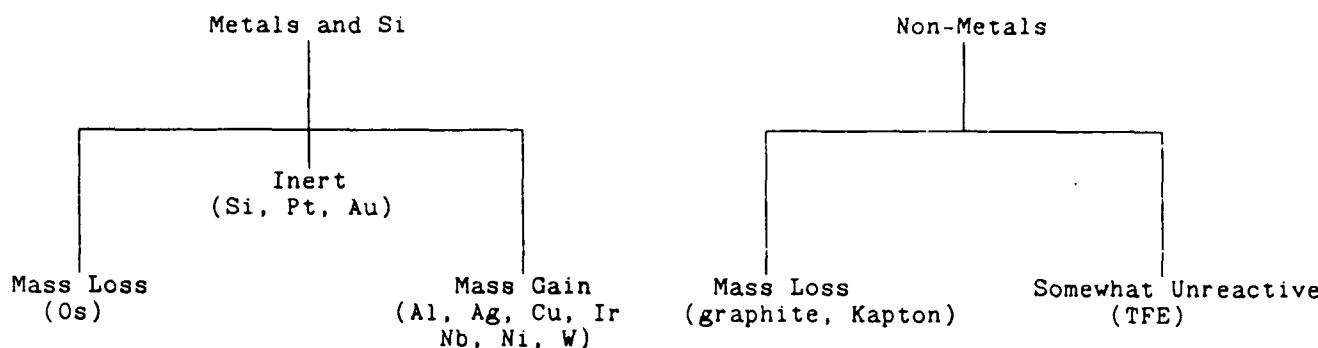


Fig. 1 Schematic representation of materials experiments.

Table 1 Composition of the low-Earth-orbit atmosphere^a

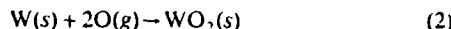
Species	Altitude, km			
	250	300	350	400
O	1.9(9) ^b	7.5(8)	3.1(8)	1.3(8)
N ₂	4.3(8)	8.6(7)	1.9(7)	4.2(6)
O ₂	2.7(7)	4.4(6)	7.6(5)	1.4(5)

^aReference 1. ^bRead $a(b)$ as $a \times 10^b$. Units are molecules/cm³.

this paper. Under equilibrium conditions, the equilibrium constant of a given reaction, k_{eq} , is given by

$$\Delta F = -RT \ln k_{eq} = \Delta H - T\Delta S \quad (1)$$

where T is in absolute degrees, ΔH is the heat of reaction, and ΔS is the entropy. In this work we will use, as units, kilocalories per mole for ΔH . In the discussion ΔH will be calculated first for $T = 298^\circ\text{K}$, since that is the temperature for which there are accurate thermochemical data. For spacecraft in low Earth orbit, the interaction between the O atoms and the surfaces occurs at orbital velocity, i.e., for O atoms having a velocity of ~ 7.7 km/s ($\equiv 113$ kcal/mole). Therefore, under these circumstances we will simply add the kinetic energy of O atoms to the heat term; i.e., we will add 102 kcal/mole per O atom to the reaction. This means that the heat term, ΔH , will be much larger than the entropy term; consequently, the entropy term will not be considered in the treatment. The concept of equilibrium is somewhat hazy under the condition of low Earth orbit, since the kinetic energy of O (~ 7.7 km/s) corresponds to a temperature of $\sim 60,000^\circ\text{K}$, and the actual temperature of the surface is $\sim 250^\circ\text{K}$. For this reason, only the heats of reaction will be calculated (i.e., k_{eq} will not be considered). The appropriateness of the treatment can be seen by considering the reaction



a reaction that will be considered in more detail later. The JANAF Tables¹¹ give $\sim 141.0 \pm 0.2$ kcal/mole and $+59.6 \pm 0.02$ kcal/mole for $\Delta H_f[\text{WO}_2, c, 298]$ and $\Delta H_f[\text{O}, g, 298]$, respectively. The ΔS at 298°K is given as 12.1, 7.8, and 38.5 cal deg⁻¹ mole⁻¹ for $\text{WO}_2(s)$, $\text{W}(s)$, and $\text{O}(g)$, respectively. If we substitute the heat and entropy of reaction (-260 kcal/mole and -72.7 cal deg⁻¹ mole⁻¹, respectively) derived from these values into Eq. (1), we obtain an equilibrium constant of $> 10^{100}$. Had we not included the entropy term, k_{eq} would have been slightly larger. Adding the kinetic energy of O to its ΔH_f leads to a still larger equilibrium constant. The implication of this discussion is that the entropy term is unimportant by comparison with heat of reaction and with orbital velocity. Changing the temperature from 298 to 200°K has a negligible effect, since $[-\Delta H/RT]$ is quite large; a change of 30% in T means that k_{eq} will increase by a factor of $\sim 6(300/200)$.

Table 2 Summary of materials experiments in space^a

Material	Thickness, nm	Effect	Thermal conductivity ^b
Al film	3.4	Oxidation	2.36
Ir film	32.3	Oxidation	1.48
Ni film	54.2	Oxidation	0.941
Nb film	16.8	Oxidation	0.533
Ag bulk	Bulk	Oxidation	4.29
Cu bulk	Bulk	Oxidation	4.03
W film	4.8	Oxidation	1.77
Os bulk	Bulk	Loss	0.90
Au film	35.5	None	3.19
Pt film	10	None	0.717
Si bulk	Bulk	None	1.68
Be			2.18
Ti			0.224

^aThe data about W film were taken from Ref. 5 and all the other data were taken from a summary by Ref. 8.

^bReference 10. Units are $\text{W} \cdot \text{cm}^{-1} \cdot \text{K}^{-1}$ at $T = 273^\circ\text{K}$.

It should be kept in mind that, for a given reaction, even if it is very exothermic, the heat of reaction may not necessarily lead to vaporization, since the heat capacity and thermal conductivity of the sample will determine how this energy is absorbed. In this vein, materials that are good thermal conductors, e.g., Al, Ag, and Cu, would not be expected to evaporate due to the quick conduction of the heats of reaction. On the other hand, materials that are poor thermal conductors would be expected to evaporate if the heat of reaction is sufficient to cause sublimation or decomposition.

The conversion of kinetic energy of a projectile (E_{lab}) into usable energy in the collision of the projectile with a target (E_{CM} , the center-of-mass energy), depends on the relative masses of the collision partners:

$$E_{\text{CM}} = [(m_p)/(m_p + m_t)] E_{\text{lab}}$$

where m_p is the mass of projectile (O) and m_t the mass of the target (surface). For the case under discussion, the collision of O atoms with a surface, the mass of the target is so much greater than that of the projectile that the $E_{\text{CM}} \sim E_{\text{lab}}$.

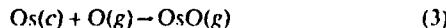
Metals and Silicon

We will discuss briefly the probable cause for loss or gain of mass in view of existing thermochemical data. The metals that are not affected by the low-Earth-orbit environment are inert in the thermochemical sense; this means that they will not be affected by the atomic energy, even though they may still participate in catalytic reactions.

Metals that Lose Mass

Os is the only metal that shows a definite mass loss; it is also the only metal that has a stable gaseous oxide at high temperatures.

Osmium. Of the metals that have been exposed to low-Earth-orbit environment, Os is the only one that has a stable gaseous oxide, OsO_4 , at room temperature. Correlations of thermochemical data¹¹ lead to $\Delta H_f[\text{OsO}_4, \text{g}, 298] = +105 \pm 20$ kcal/mole. This means that at 298°K

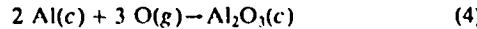


is endothermic by 46 kcal/mole. If all of the orbital energy is available for reaction, then reaction (3) can become exothermic by 67 kcal/mole in the low-Earth-orbit environment.

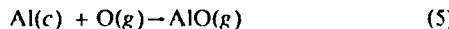
Metals that Gain Mass

Even though these metals may have gaseous monoxides that are stable at high temperatures, their solid oxides are extremely stable.

Aluminum. $\text{Al}_2\text{O}_3(\text{c})$ is extremely stable, with a heat of formulation of $\Delta H_f[\text{Al}_2\text{O}_3, \text{c}, 298] = -400.5 \pm 0.3$ kcal/mole.¹¹ This means that at 298°K the reaction

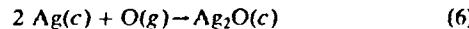


is exothermic by 578 kcal/mole. At orbital energies it is even more exothermic. The fact that Al exposed to the low-Earth-orbit environment does not lose mass suggests that

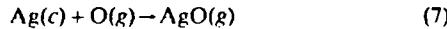


does not occur, even though it is simpler than reaction (4) and even though it is exothermic in the low Earth environment by 43 kcal/mole at 298°K [$\Delta H_f[\text{AlO}, \text{g}, 298] = 16 \pm 2$ kcal/mole].¹² This observation of Al also illustrates the difficulty of relating the high impact velocity of O to a thermodynamic temperature, since the vaporization of $\text{Al}_2\text{O}_3(\text{c})$ is known to yield other gaseous oxides.¹³

Silver. The gain of mass by silver films when exposed to the environment of low Earth orbit is undoubtedly due to the formation of $\text{Ag}_2\text{O}(\text{c})$, since $\Delta H_f[\text{Ag}_2\text{O}, \text{c}, 298] = -7.4$ kcal/mole, which means that

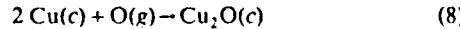


is exothermic at room temperature by 67 kcal/mole. At orbital velocity, the reaction is even more exothermic. Again, the fact that mass is not lost means that

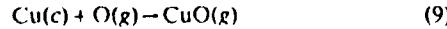


does not occur, even though in the low-Earth-orbit environment it would be exothermic by 97 ± 5 kcal/mole, since $\Delta H_f[\text{AgO}, \text{g}, 298] = 75 \pm 5$ kcal/mole.¹² It may be that the high thermal conductivity of Ag is sufficient to dissipate the heat of reaction before it can be distributed to the reactants. Incidentally, it is worth noting that reaction (7) is the basis of an attempt to measure the height profile of atomic oxygen in the upper atmosphere¹⁴; in this experiment, however, the velocity of atomic oxygen is essentially thermal. The reaction alters the conductivity of a thin film of Ag by an amount that is proportional to the integrated flux of atomic oxygen.

Copper. $\Delta H_f[\text{Cu}_2\text{O}, \text{c}, 298] = -40.3$ kcal/mole,¹⁵ leading to the conclusion that

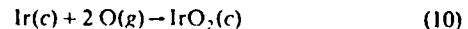


is exothermic by 99 kcal/mole. It is even more so if orbital energy of O is taken into account. $\Delta H_f[\text{CuO}, \text{g}, 298]$ is 76 ± 5 kcal/mole,¹² which again means that



can, but does not, take place in the low Earth environment. Part of the reason may be due to the efficient conduction of the heat of reaction by the bulk.

Iridium. Iridium has a stable gaseous monoxide at high temperatures, but not at room temperature. Ir has a stable, condensed oxide, $\text{IrO}_2(\text{c})$, whose heat of formation, $\Delta H_f[\text{IrO}_2, \text{c}, 298]$, is reported to be -65.6 kcal/mole. This means that

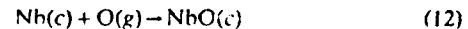


is exothermic by 183.6 kcal/mole. Since Ir is a fairly good thermal conductor, it is apparent that all of this heat is transferred to the substrate. It will be useful in future experiments to measure the gaseous composition above a surface of Ir exposed to a low-Earth-orbit environment. The stable gaseous oxide of Ir is $\text{IrO}(\text{g})$. Recent analysis¹² leads to $\Delta H_f[\text{IrO}, \text{g}, 298] = +121 \pm 10$ kcal/mole, which means that the reaction leading to the formation of gaseous IrO ,

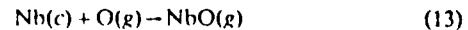


is endothermic by 62 kcal/mole at 298°K .

Niobium. $\Delta H_f[\text{NbO}, \text{c}, 298] = -97$ kcal/mole,¹⁵ leading to a heat of reaction of -156 kcal/mole for

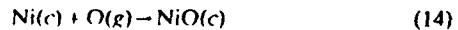


at 298°K . Since the orbital velocity leads to O atoms having 102 kcal/mole of kinetic energy, the reaction in low-Earth-orbit becomes exothermic by 269 kcal/mole. As in the case of Ni, the reaction that might lead to mass loss,

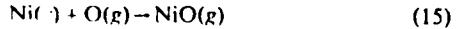


does not seem to take place, even though at orbital velocities it would be exothermic by 122 kcal/mole, since $\Delta H_f[\text{NbO}, \text{g}, 298] = +51$ kcal/mole.¹⁵

Nickel. $\Delta H_f[\text{NiO}, \text{c}, 298] = -57.3$ kcal/mole,¹⁵ leading to a heat of reaction of ~ -117 kcal/mole for

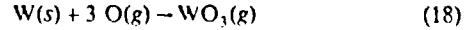
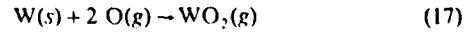
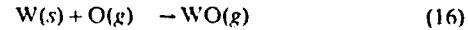
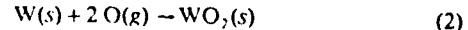


at 298°K . If the kinetic energy of O is taken into account, reaction (14) would be exothermic by 230 kcal/mole. In contrast, that could lead to mass loss,



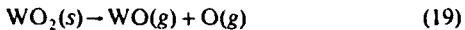
is exothermic by -97 kcal/mole at orbital altitudes, since $\Delta H_f[\text{NiO}, \text{g}, 298] = -5$ kcal/mole.¹⁵

Tungsten. Tungsten has a number of oxides stable at high temperatures, $\text{WO}(\text{g})$, $\text{WO}_2(\text{g})$, $\text{WO}_2(\text{s})$, and $\text{WO}_3(\text{g})$; however, the gaseous oxides are not stable at room temperature. Consequently, the processes by which these oxides can be formed,



are not observed at room temperature. Recent analyses of thermochemical data give $\Delta H_f[\text{WO}, \text{g}, 298] = +102 \pm 10$ kcal/mole,¹² $\Delta H_f[\text{WO}_2, \text{g}, 298] = +18.3 \pm 2$ kcal/mole,¹¹ $\Delta H_f[\text{WO}_2, \text{s}, 298] = -140.9 \pm 0.2$ kcal/mole,¹¹ and $\Delta H_f[\text{WO}_3, \text{g}, 298] = -70 \pm 7$ kcal/mole.¹¹ From these heats of formation and from $\Delta H_f[\text{O}, \text{g}, 298] = +59.6 \pm 0.02$ kcal/mole, we derive at 298°K ΔH_f [Eq. (16)] = $+42 \pm 10$

kcal/mole, ΔH_f [Eq. (17)] = -101 ± 2 kcal/mole, ΔH_f [Eq. (2)] = -260 ± 1 kcal/mole, and ΔH_f [Eq. (18)] = -248.8 ± 7 kcal/mole. At orbital velocity, the heats of reaction become ΔH_f [Eq. (16)] = -71 kcal/mole, ΔH_f [Eq. (17)] = -214 ± 2 kcal/mole, ΔH_f [Eq. (18)] = -475 ± 1 kcal/mole, and ΔH_f [Eq. (18)] = -566 ± 7 kcal/mole; in other words, they all become exothermic. The fact that mass gain is observed suggests that $\text{WO}_2(s)$ is formed. Since mass gain is observed, we have to conclude that the combination of thermal capacity of the substrate, the thermal conductivity of the sample, and internal excitation are enough to stabilize the solid oxide phase so that it will not be vaporized via



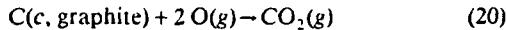
which is endothermic by 301 kcal/mole at 298°K. It would be interesting in this connection to look for thermal emission from a tungsten surface exposed to low-Earth-orbit environment.

Inert Materials

Three materials, Au, Pt, and Si, showed no effect from exposure to low-Earth-orbit environment. Au and Pt are easily understood since they do not have stable oxides in either the gas phase or in the condensed phase at room temperature or at high temperatures. Si does have a stable oxide, SiO_2 (silica), which was not formed even though its formation is favored thermochemically. This seeming inertness is probably due to the formation of a stable layer of $\text{SiO}(c)$ (actually $\text{Si}-\text{SiO}_2$, which is a stable condensed phase).

Carbonaceous Materials

The situation in this case is somewhat complicated by the lack of thermochemical information regarding many of the materials that have been tested. However, for C (in the form of graphite) we can make estimates of heats of reaction for



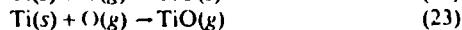
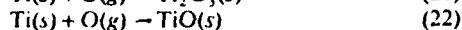
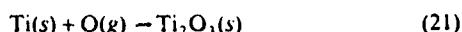
using $\Delta H_f[\text{CO}_2, g, 298] = -94.1$ kcal/mole¹⁵ and $\Delta H_f[\text{C, graphite}, 298] = 0$ kcal/mole,¹⁵ we derive ΔH_f [Eq. (20)] = -153 , exclusive of the kinetic energy of O atoms. It is likely that the other composites that have been tested will have similar thermodynamic properties; i.e., they will be eroded rather quickly. Recently reported work has extended these studies to Kapton and other insulators that are found to be affected rather quickly upon exposure to atomic oxygen in low Earth orbit.¹⁶

Applications

In view of the treatment mentioned earlier, we consider below two materials that have not been flown in low-Earth-orbit erosion experiments: Ti and Be. Both are candidates for use in space because of special properties.

Titanium

Even though no experiments have been conducted on Ti, it is likely to be a candidate for space structures because of its high strength and low density. Titanium is a relatively poor thermal conductor; thus we might expect that any heat generated in the reaction of the Ti with atomic oxygen will go into the sample rather than being conducted into the bulk. We consider in this analysis three possible reactions:



Using $\Delta H_f[\text{Ti}_2\text{O}_3, s] = -363 \pm 2$ kcal/mole, $\Delta H_f[\text{TiO}, s, 129.7 \pm 1]$ kcal/mole, and $\Delta H_f[\text{TiO}, g, 298] = 13 \pm 2$

kcal/mole,¹¹ we derive ΔH_f [Eq. (21)] = -540 ± 5 kcal/mole, ΔH_f [Eq. (22)] = -188 ± 3 kcal/mole, and ΔH_f [Eq. (23)] = -46 ± 3 kcal/mole, all at 298°K. The heat of vaporization of $\text{TiO}(s)$, $\Delta H_f[\text{TiO}, s, 298]$, is given by the difference in the

Table 3 Summary of temperatures of formation used in this paper

Species	$\Delta H_f(298^\circ\text{K, thermal}),$ kcal/mole	$S(298^\circ\text{K}),$ cal deg ⁻¹ mole ⁻¹
Ag(c)	0	10.2 ^a
AgO(g)	+75 ± 5 ^b	
Ag ₂ O(c)	-7.4 ^a	29.0
Al(c)	0	6.76
AlO(g)	+16 ± 2 ^c	52.19
Al ₂ O ₃ (c)	-400.5 ± 0.3 ^c	12.18
BeO(c)	-145.4 ± 0.9 ^c	3.29
BeO(g)	+32.6 ^c	47.23
C(graphite)	0	1.37
C(diamond)	+0.5 ^a	0.56
CO(g)	-27.2 ^a	31.69
CO ₂ (g)	-94.1 ± 0.01 ^c	51.10
Cu(c)	0	7.93
CuO(g)	+73.2 ± 10 ^c	56.1
Cu ₂ O(c)	-40.8 ± 0.5 ^c	22.1
Ir(c)	0	8.5 ^a
IrO(g)	+121 ± 10 ^b	
IrO ₂ (c)	-274.1 ^a	
Nb(c)	0	8.71
NbO(c)	-100.3 ± 3 ^c	[11 ± 2] ^c
NbO(g)	+47.5 ± 5 ^c	57.12
Ni(c)	0	7.14 ^c
NiO(c)	-57.3 ^a	9.01
NiO(g)	+75 ^a	
O(g)	+59.55 ± 0.02 ^c	38.49
Os(c)	0	7.8 ^a
OsO(g)	+105 ± 20 ^b	
Si(c)	0	4.49
SiO(g)	-24.0 ± 2 ^c	50.56
SiO ₂ (c)	-217.7 ^a	10.0
Ti(c)	0	7.35
TiO(c)	-129.7 ± 3 ^c	8.31
TiO(g)	+13 ± 2 ^c	55.80
Ti ₂ O ₃ (s)	-363 ± 2 ^c	18.46
W(c)	0	7.81
WO(g)	+102 ± 10 ^c	[59.3] ^c
WO ₂ (c)	-140.9 ± 0.2 ^c	12.08
WO ₂ (g)	+18.3 ± 7 ^c	[68.3] ^c
WO ₃ (c)	-201.4 ± 0.2 ^c	18.14
WO ₃ (g)	-70 ± 7 ^c	[68.5] ^c

^aReference 15. ^bReference 12. ^cReference 11.

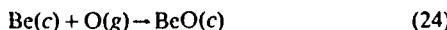
Table 4 Summary of reactions and their temperatures

Reaction	$\Delta H_f(298^\circ\text{K}),$ kcal/mole
(2) $\text{W}(s) + 2 \text{O}(g) \rightarrow \text{WO}_2(s)$	-260 ± 1
(3) $\text{Os}(c) + \text{O}(g) \rightarrow \text{OsO}(g)$	+46 ± 20
(4) $2 \text{Al}(c) + 3 \text{O}(g) \rightarrow \text{Al}_2\text{O}_3(c)$	-578 ± 1
(5) $\text{Al}(c) + \text{O}(g) \rightarrow \text{AlO}(g)$	-43 ± 2
(6) $2 \text{Ag}(c) + \text{O}(g) \rightarrow \text{Ag}_2\text{O}(c)$	-67
(7) $\text{Ag}(c) + \text{O}(g) \rightarrow \text{AgO}(g)$	+15
(8) $2 \text{Cu}(c) + \text{O}(g) \rightarrow \text{Cu}_2\text{O}(c)$	-99
(9) $\text{Cu}(c) + \text{O}(g) \rightarrow \text{CuO}(g)$	+16
(10) $\text{Ir}(c) + \text{O}(g) \rightarrow \text{IrO}(g)$	+62 ± 10
(11) $\text{Ir}(c) + \text{O}(g) \rightarrow \text{IrO}_2(g)$	+62 ± 10
(12) $\text{Nb}(c) + \text{O}(g) \rightarrow \text{NbO}(c)$	-156
(13) $\text{Nb}(c) + \text{O}(g) \rightarrow \text{NbO}_2(g)$	-9
(14) $\text{Ni}(c) + \text{O}(g) \rightarrow \text{NiO}(c)$	-117
(15) $\text{Ni}(c) + \text{O}(g) \rightarrow \text{NiO}_2(g)$	+16
(16) $\text{W}(s) + \text{O}(g) \rightarrow \text{WO}(g)$	-42 ± 10
(17) $2 \text{W}(s) + 2 \text{O}(g) \rightarrow \text{WO}_2(g)$	-102 ± 2
(18) $\text{W}(s) + 3 \text{O}(g) \rightarrow \text{WO}_3(g)$	-249 ± 7
(19) $\text{WO}_2(s) \rightarrow \text{WO}(g) + \text{O}(g)$	+301 ± 10
(20) $\text{C}(c, \text{graphite}) + 2\text{O}(g) \rightarrow \text{CO}_2(g)$	-153
(21) $\text{Ti}(s) + \text{O}(g) \rightarrow \text{Ti}_2\text{O}_3(s)$	-540
(22) $\text{Ti}(s) + \text{O}(g) \rightarrow \text{TiO}(s)$	-188
(23) $\text{Ti}(s) + \text{O}(g) \rightarrow \text{TiO}(g)$	-46
(24) $\text{Be}(c) + \text{O}(g) \rightarrow \text{BeO}(g)$	-305

heat of formation of the solid and the gas phases of TiO and is equal to $+142.7 \pm 2$ kcal/mole. Oxidation of the surface via reactions (21) and (22) might lead to the formation of an insulating layer, but since the heat of vaporization of $\text{TiO}(s)$ formed in reaction (22) is less than the heat of reaction at orbital velocity, it seems likely that any oxide formed will evaporate. Likewise, the formation of $\text{Ti}_2\text{O}_3(s)$ is likely to be temporary because of the large amounts of energy deposited in the interaction at orbital velocities; thermochemical measurements indicate that all of the stable solid phases of $\text{Ti}-\text{O}$ vaporize to give $\text{TiO}(g)$ and $\text{Ti}(g)$.¹⁷ The combination of the thermochemical data with the thermal conductivity information shown in Table 2 suggests that a Ti surface will be eroded on exposure to low Earth orbit.

Beryllium

Be has not been flown on the Space Shuttle; however, it is a strong candidate for use in optical instrumentation in low Earth orbit because of its light weight and optical properties. Be is a relatively good thermal conductor, as the data of Table 2 shows. As in the case of Ti , we consider the following reaction:



$\Delta H_f[\text{Be}, c, 298] = 0$, whereas $\Delta H_f[\text{BeO}, c, 298] = -145.4 \pm 0.9$ kcal/mole.¹¹ Reaction (24) is exothermic by 205 kcal/mole at 298°K. At 7.7 km/s the reaction becomes exothermic by 318 kcal/mole at 298°K. Even though the heat of vaporization of $\text{BeO}(c)$ [$\Delta H_v[\text{BeO}, c, 298] = 178 \pm 3$ kcal/mole¹¹] suggests that for Be mass loss should occur; however, because of its good thermal conductivity, it is likely that the interaction would lead to the buildup of an oxide coating at low Earth orbit (i.e., where O is a dominant constituent), at least for bulk samples of Be . On the other hand, at higher altitudes where $[\text{H}]$ becomes comparable with $[\text{O}]$, the interaction may be quite benign in that no changes will take place.

Conclusions

This paper has attempted to present a discussion of the thermochemical considerations that play a role in determining the fate of samples exposed to low-Earth-orbit environment. Tables 3 and 4 present summaries of the temperatures of formation as well as the equilibria used in this analysis. The thermochemical data were obtained from standard compilations, and the reactions used in the study are *simple* gas-solid reactions. The *simplicity* of the reactions is needed because of the rarified atmosphere in the low-Earth-orbit environment (thus arguing against multiple collisions). The discussion presented in this paper leads to three implications: 1) erosion of materials, 2) buildup of materials and change in the physical properties of materials, and 3) effects of these interactions on surfaces and sensors on low-Earth-orbit stations. These effects will be discussed briefly below.

Erosion of Materials

The erosion of materials leads to weakening of structures in space, leading to profound implication for the integrity of long-lived structures in space. For example, in some of the experiments, it was observed that Os coated with Pt still showed some erosion near the edges of the samples.⁷ Leger and Visentine¹⁸ have estimated that the Space Station may not last a full solar cycle if the designs and the materials being contemplated are not changed. In addition, mirrors and other large optical structures that need to be made of very light ultralow expansion material are liable to be modified drastically unless the designs include allowances for these contingencies.

Buildup of Materials

The buildup of materials, through the formation of oxides, leads to a change in the electrical properties of a metal. For ex-

ample, the formation of $\text{NiO}(c)$, $\text{Cu}_2\text{O}(c)$, and $\text{Al}_2\text{O}_3(c)$ on the respective metals can lead to electrical breakdowns and to different thermal properties; the transformation of metals into oxides will have profound effects on optical properties of surfaces that incorporate these metals in their structures, as in the case of mirrors and reflectors being contemplated for space use. The possible buildup of an oxide coating on a Ti surface means that the surface will be insulated and may eventually result in charging and arcing of the surfaces.

Effect on Optical Experiments

Mass loss from a surface exposed to the low-Earth-orbit environment means that either a gaseous product (chemical reaction) or a solid particulate (something like sputtering) has formed. Because of the high flux of O atoms ($\text{flux} = [\text{O}]^*v$), the column density of any gaseous products formed at the surface is likely to be equally high. Sensors and optical surfaces are affected because their surfaces become rough quickly, leading to degraded performance; in addition, if there is a gaseous cloud in front of the optical instrument, either extraneous signals may be generated by the cloud (via fluorescence) or the performance of the sensor may be degraded (via absorption of radiation by the cloud). For example, an organic surface near an infrared sensor may generate enough CO_2 via reactions (20) and (21) to raise the threshold in the $4.3\text{-}\mu$ region. Likewise, the presence of Ti or W in that region may generate enough $\text{WO}(g)$ and $\text{TiO}(g)$ to distort observations in the 10 - to $12\text{-}\mu$ region, where the fundamentals seem to cluster.¹⁹

To obtain a more definitive analysis of the problem, a number of diverse studies are needed:

- 1) More quantitative space experiments.
- 2) Repetition of the data in Table 2 under more controlled conditions (e.g., all bulk sample or all film samples) in order to be able to separate the effect of heat capacity and thermal conductivity from the reactivity of the metal.
- 3) Better thermochemical data, particularly for the oxides of W , Ir , and Os .
- 4) Kinetic studies of the reaction of fast O atoms with surfaces.
- 5) Accurate surface science experiments: characterization of the surface before and after exposure to low-Earth-orbit environment.
- 6) Analysis of the gaseous composition above surfaces exposed to low-Earth-orbit environment.

Acknowledgment

I would like to thank one of the reviewers for helpful comments about the interpretation of the data given in Ref. 7.

References

- ¹Jursa, A. S. (ed.), *Handbook of Geophysics and the Space Environment*, National Technical Information Service, Springfield, VA, 1985.
- ²Green, B. D., Caledonia, G. E., and Wilkerson, T. D., "The Shuttle Environment: Gases, Particulates, and Glow," *Journal of Spacecraft and Rockets*, Vol. 22, Sept.-Oct. 1985, pp. 500-511.
- ³Kofsky, I. L. and Barrett, J. L., "Spacecraft Surface Glows," *Nuclear Instruments & Methods*, Vol. B14, April 1986, pp. 480-486.
- ⁴Murad, E., "Glow of Spacecraft in Low Earth Orbit," *Physics of Space Plasmas* (1985-1987), Vol. 6, edited by T. Chang, J. Belcher, J. R. Jasperse, and B. G. Grew, Scientific Publishers, Cambridge, MA, 1987, p. 147.
- ⁵Garrett, H. B., Chutjian, A., and Gabriel, S., "Space Vehicle Glow-A Review," *Journal of Spacecraft and Rockets*, Vol. 25, Sept.-Oct. 1988, pp. 321-340.
- ⁶Peters, P. N., Linton, R. C., and Miller, E. R., "Results of Apparent Atomic Oxygen Reactions on Ag , C , and Os Exposed During the Shuttle STS-4 Orbits," *Geophysical Research Letters*, Vol. 10, July 1983, pp. 569-571.
- ⁷Peters, P. N., Gregory, J. C., and Swann, J. T., "Effects on Optical System from Interactions with Oxygen Atoms in Low Earth Orbit," *Applied Optics*, Vol. 25, April 1986, pp. 1290-1298.

⁸Leger, L. J., Santos-Mason, B., Visentine, J., and Kuminescu, J., "Review of LEO Flight Experiments," *Proceedings of the NASA Workshop on Atomic Oxygen Effects*, edited by D. E. Brinza, Jet Propulsion Lab., Pasadena, CA, Pub. 87-14, 1987, pp. 1-10.

⁹Gregory, J. C., "Interaction of Hyperthermal Atoms on Surfaces in Orbit: The University of Alabama Experiment," *Proceedings of the NASA Workshop on Atomic Oxygen Effects*, edited by D. E. Brinza, Jet Propulsion Lab., Pasadena, CA, Pub. 87-14, 1987, pp. 29-36.

¹⁰Ho, C. Y., Powell, R. W., and Liley, P. E., "Thermal Conductivity of the Elements," *Journal of Physical and Chemical Reference Data*, Vol. 1, March 1972, pp. 279-421.

¹¹American Chemical Society, *JANAF Thermochemical Tables*, 3rd. ed., American Chemical Society, Washington, DC, 1986.

¹²Pedley, J. B., and Marshall, E. M., "Thermochemical Data for Gaseous Monoxides," *Journal of Physical and Chemical Reference Data*, Vol. 12, Dec. 1983, pp. 967-1025.

¹³Lamoreaux, R. H., Hildenbrand, D. L., and Brewer, L., "High Temperature Vaporization Behavior of Oxides. II. Oxides of Be, Mg, Ca, Sr, Ba, B, Al, Ga, In, Tl, Sn, Pb, Zn, Cd, and Hg," *Journal of Physical and Chemical Reference Data*, Vol. 16, Sept. 1987, pp. 419-445.

¹⁴Henderson, W. R. and H. I. Schiff, "A Simpler Sensor for the Measurement of Oxygen Height Profiles in the Upper Atmosphere,"

Planetary and Space Science, Vol. 18, 1970, pp. 1527-1534.

¹⁵Wagman, D. D., Evans, W. H., Parker, V. B., Schumm, R. H., Halow, I., Bailey, S. M., Churney, K. L., and Nutall, R. L., "The NBS Tables of Chemical Thermodynamic Properties: Selected Values for Inorganic and C₁ and C₂ Organic Substances in SI Units," *Journal of Physical and Chemical Reference Data*, Vol. 11, Suppl. 2, June 1982, pp. 1-392.

¹⁶Zimcik, D. G. and Maag, C. R., "Results of Apparent Atomic Oxygen Reactions with Spacecraft Materials During Shuttle Flight STS-41G," *Journal of Spacecraft and Rockets*, Vol. 25, March-April 1988, pp. 162-168.

¹⁷Gilles, P. W., Carlson, K. D., Franzen, H. F., and Wahlbeck, P. G., "High-Temperature Vaporization and Thermodynamics of the Titanium Oxides. I. Vaporization Characteristics of the Crystalline Phases," *Journal of Chemical Physics*, Vol. 46, April 1967, pp. 2461-2465.

¹⁸Leger, L. J. and Visentine, J. T., "A Consideration of Atomic Oxygen Interactions with the Space Station," *Journal of Spacecraft and Rockets*, Vol. 23, Sept.-Oct. 1986, pp. 505-511.

¹⁹Huber, K. P. and Herzberg, G., *Molecular Spectra and Molecular Structure: IV. Constants of Diatomic Molecules*, Van Nostrand, New York, 1979.

**Recommended Reading from the AIAA
Progress in Astronautics and Aeronautics Series . . .**



**Monitoring Earth's Ocean, Land
and Atmosphere from Space:
Sensors, Systems, and Applications**

Abraham Schnapf, editor

This comprehensive survey presents previously unpublished material on past, present, and future remote-sensing projects throughout the world. Chapters examine technical and other aspects of seminal satellite projects, such as Tiros/NOAA, NIMBUS, DMS, LANDSAT, Seasat, TOPEX, and GEOSAT, and remote-sensing programs from other countries. The book offers analysis of future NOAA requirements, spaceborne active laser sensors, and multidisciplinary Earth observation from space platforms.

TO ORDER: Write AIAA Order Department,
370 L'Enfant Promenade, S.W., Washington, DC 20024
Please include postage and handling fee of \$4.50 with all
orders. California and D.C. residents must add 6% sales
tax. All foreign orders must be prepaid.

1985 830 pp., illus. Hardback
ISBN 0-915928-98-1
AIAA Members \$59.95
Nonmembers \$99.95
Order Number V-97